

MARK SCHEME for the October/November 2012 series

9701 CHEMISTRY

9701/33

Paper 3 (Advanced Practical Skills 1),
maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2012 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

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Question	Sections	Indicative material	Mark
1 (a)	PDO Recording	I All columns correctly headed and correct units given for all columns except for rate/(1000/time) e.g. /s, (s), time in s, time in seconds.	1
		II Records all times to the nearest second. Allow for only 5 expts carried out.	1
	MMO Decisions	III Additional experiment (experiment 6) uses volume of FA 1 $\geq 3 \text{ cm}^3$ of any other, and adds water to make 50 cm^3 . Other volumes are those specified.	1
	PDO Display	IV Candidate gives all values of (1000/time) to 3 sig fig – ignore calculation or rounding errors (minimum of 4 expts carried out).	1
	ACE Interpretation	V All values of (1000/time) correctly calculated to sig fig shown by candidate (minimum of 4 expts carried out).	1
	MMO Quality	VI and Experiments 2 and 4: VII calculate $100(2t_2 - t_4)/t_4 \leq 20\%$ for 1 mark; $\leq 10\%$ for 2 marks.	6
		VIII and Experiments 3 and 5: IX calculate $100(3t_3 - t_5)/t_5 \leq 30\%$ for 1 mark; $\leq 10\%$ for 2 marks.	
		X and Experiments 4 and 5: XI calculate $100(2t_4 - t_5)/t_5 \leq 30\%$ for 1 mark; $\leq 10\%$ for 2 marks.	
		If the candidate has not completed the 5 th experiment, marks VI and VII are available. Also check Experiments 1 and 2: t_2 should equal to $t_1 \times 5/4$. Use the 10% and 20% boundaries.	
		If only the first three experiments are completed, award Q marks based on Experiments 1 and 2 (as above).	
	<i>(If 50, 45, 40, 35, marks X and XI not available. Use 40 and 20 if there + 'rescue' pair as above.)</i>		
	The Examiner is to round all reaction times to the nearest second before awarding accuracy marks. <i>(Volumes FA 1/expt no as specified in Qn)</i>		
		[11]	

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(b)	PDO Layout	I Plots rate or (1000/time) on y-axis and volume of FA 1/FA 1 cm ³ on x-axis. Axes correctly labelled.	1	[5]
		II Uniform scales selected. Each scale starts at zero and highest point plotted on each axis has used more than half of the available grid.	1	
		III and IV Examiner to check all plotted points. IV Points must be correct to ½ small square and in correct small square.	2	
		V Draws a “best-fit” straight line – one that passes close to the majority of points and points are balanced. The line does not have to pass through the origin. (Allow curve if appropriate.)	1	
(c)	ACE Conclusions	Depth (of solution) is greater, ... so time is shorter/less//time is faster//fewer seconds (time is conditional on depth)	1	[2]
		or solution/liquid depth unchanged so reaction time unchanged for 1 mark.	1	
(d)	ACE Interpretation	Give one mark for a concentration of 0.021/0.0214/0.02143 mol dm ⁻³ for expt 5.	1	[2]
	PDO Display	Working shown must include correct use of 70.	1	

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(e)	ACE Interpretation	Two pieces of evidence with no conclusion or one piece and conclusion. 2 nd piece of evidence and conclusion.	1	
		<p>Evidence for 'correct'</p> <p>(i) a straight line/(line with) constant gradient (ii) straight line passes through origin (if appropriate from results) is 2 pieces of evidence (iii) line passes through origin = 1 if line drawn is straight</p> <p>Evidence for 'incorrect'</p> <p>(i) a curve has been drawn/no straight line/not constant gradient (ii) straight line does not pass through the origin (iii) points too scattered/not on best fit line (iv) a curve drawn but expect straight line = 2</p> <p><i>A straight line, not passing through the origin could score both marks depending on explanation given (proportional but not directly proportional). If two points are compared they must be on or very close to the graph line.</i></p>	1	[2]
(f)	ACE Interpretation	Candidate correctly evaluates each % uncertainty.	1	[1]
(g)	ACE Improvement/s	Constant volume of FA 1 .	1	
		Varies volume of FA 2 and water correspondingly (Volume FA 2 + H ₂ O same).	1	[2]
Total			25	

FA 3 is $\text{CuCl}_2(\text{aq})$; FA 4 is $\text{AlK}(\text{SO}_4)_2(\text{aq}) + \text{KI}(\text{aq})$; FA 5 is $\text{FeCl}_3(\text{aq})$; FA 6 is $\text{Pb}(\text{NO}_3)_2$

2	(a)	MMO Collection	Records a blue / greenish-blue ppt / solid with FA 3 and Na_2CO_3 .	1	[4]
			Records a brown / rust / orange-brown / red-brown ppt / solid with FA 5 and Na_2CO_3 .	1	
			Records effervescence with FA 5 (or FA 3).	1	
		MMO Decisions	Tests <u>gas</u> evolved with limewater. Allow from effervescence.	1	
	(b)	MMO Collection	Records a white precipitate with silver nitrate solution and soluble in aqueous ammonia.	1	[1]
	(c)	MMO Collection	Records yellow-brown / orange-brown / brown / tan colour (solid/solution) (formed on mixing FA 4 and FA 3). Allow dark brown for solution only . Allow (qualified) brown solution with white / off-white / grey ppt. Dark / deep blue / blue-black / black / purple colour on adding starch solution	1 1	[2]
	(d)	MMO Collection	Mark the observations in the table horizontally or vertically to maximise marks available to the candidate.	4	[4]

Test	Observations			
	FA 3	FA 4	FA 5	FA 6
NaOH(aq)	blue ppt not dark/deep blue ppt	white ppt (which dissolves as more added/then dissolves)	red-brown / orange-brown / brown / rust ppt (not dark / deep brown)	white ppt
excess NaOH	ppt insoluble (no change no observation provided ppt above)	ppt soluble (if no ppt in 1 st box allow no change)	ppt insoluble (no change no observation provided ppt above)	ppt soluble (not no change after 'no ppt')
NH ₃ (aq)	blue ppt not dark/deep blue ppt	white ppt	red-brown / orange-brown / brown / rust ppt (not dark / deep brown)	white ppt
excess ammonia	(ppt soluble) deep blue soln	ppt insoluble (no change no observation provided ppt above)	ppt insoluble (no change no observation provided ppt above)	ppt insoluble (no change no observation provided ppt above)

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FA 3 is $\text{CuCl}_2(\text{aq})$; FA 4 is $\text{Al}(\text{SO}_4)_2(\text{aq}) + \text{KI}(\text{aq})$; FA 5 is $\text{FeCl}_3(\text{aq})$; FA 6 is $\text{Pb}(\text{NO}_3)_2$

(e)	ACE Conclusions	Con2	Give one mark for FA 3 Cu^{2+} /copper/copper(II) and FA 5 Fe^{3+} /iron(III).	1	[2]
		Con2	Give one mark for FA 4 and FA 6 Al^{3+} /aluminium, Pb^{2+} /lead Allow FA 4 $\text{Al}^{3+}(\text{Pb}^{2+})$ and FA 6 Al^{3+} , Pb^{2+} (There must be some correct evidence for Cu^{2+} and Fe^{3+} in (d) but does not have to be fully correct.)	1	
(f)	MMO Decisions	De7	Selects appropriate reagent to distinguish between Al^{3+} and Pb^{2+} e.g. KI, K_2CrO_4 , H_2SO_4 , HCl (not BaCl_2).	1	[1]
(g)	ACE Conclusions	Con2	No error carried forward in this section. Award the mark for: FA 3 chloride FA 4 iodide FA 5 insufficient tests	1	[1]
				Total	15